

## AP Chemistry Polyatomic Ions

#	Polyatomic Ion and Oxidation number	Name
1.	$\text{ClO}^{-1}$	Hypochlorite
2.	$\text{ClO}_2^{-1}$	Chlorite
3.	$\text{ClO}_3^{-1}$	Chlorate
4.	$\text{ClO}_4^{-1}$	Perchlorate
5.	$\text{C}_2\text{H}_3\text{O}_2^{-1}$	Acetate
6.	$\text{OH}^{-1}$	Hydroxide
7.	$\text{CN}^{-1}$	Cyanide
8.	$\text{NO}_2^{-1}$	Nitrite
9.	$\text{NO}_3^{-1}$	Nitrate
10.	$\text{HCO}_3^{-1}$	Hydrogen carbonate or Bicarbonate
11.	$\text{HSO}_4^{-1}$	Hydrogen sulfate or bisulfate
12.	$\text{HSO}_3^{-1}$	Hydrogen sulfite or bisulfite
13.	$\text{H}_2\text{PO}_4^{-1}$	Dihydrogen phosphate
14.	$\text{SCN}^{-1}$	Thiocyanate
15.	$\text{IO}_3^{-1}$	Iodate
16.	$\text{OCN}^{-1}$	Cyanate
17.	$\text{MnO}_4^{-1}$	Permanganate
18.	$\text{SO}_4^{-2}$	Sulfate
19.	$\text{SO}_3^{-2}$	Sulfite
20.	$\text{CO}_3^{-2}$	Carbonate
21.	$\text{CrO}_4^{-2}$	Chromate
22.	$\text{Cr}_2\text{O}_7^{-2}$	Dichromate
23.	$\text{SiO}_3^{-2}$	Silicate
24.	$\text{C}_2\text{O}_4^{-2}$	Oxalate
25.	$\text{HPO}_4^{-2}$	Hydrogen Phosphate
26.	$\text{O}_2^{-2}$	Peroxide
27.	$\text{PO}_4^{-3}$	Phosphate
28.	$\text{PO}_3^{-3}$	Phosphite
28.	$\text{NH}_4^{+1}$	Ammonium