**Bonding Unit: Lewis dot diagrams**

**Homework due Friday, November 14**

1. Draw Lewis structures for the following atoms, ions and compounds.

2. For the compounds: Write down the electronegativity difference and use it to identify the bond type (ionic, polar covalent, nonpolar covalent). The first one is done for you.

|  |  |  |
| --- | --- | --- |
| Barium atom | Sodium atom | Carbon atom |
| Barium ion | Iodine ion | Sulfur ion |
| HF   1. [https://encrypted-tbn0.gstatic.com/images?q=tbn:ANd9GcRRsjKN7v7EdYSRWtHe3p_6LQTTDhseNrvuvz_0fAzMQF3IwpUr](http://www.google.com/url?sa=i&rct=j&q=&esrc=s&frm=1&source=images&cd=&cad=rja&uact=8&ved=0CAcQjRw&url=http%3A%2F%2Fpixgood.com%2Fhf-lewis-dot-structure.html&ei=CMdiVPeUI8KsyQTIzoDIAQ&bvm=bv.79189006,d.aWw&psig=AFQjCNHs_3hdsEHM8uWoYWmhP_pBtXUSmw&ust=1415845994311859) | H2O | NH3 |
| EN diff= 1.8 (polar covalent) |  |  |
| LiBr | CH4 | LiBr |
|  |  |  |
| H2S | CS2 | CO2 |
|  |  |  |