Stations: Types of Chemical Reactions and Solution Stoichiometry (ch. 4)

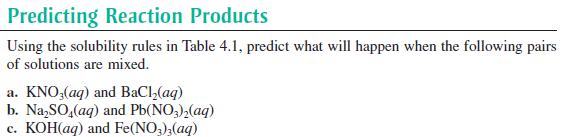
# Precipitations Reactions

NOTES

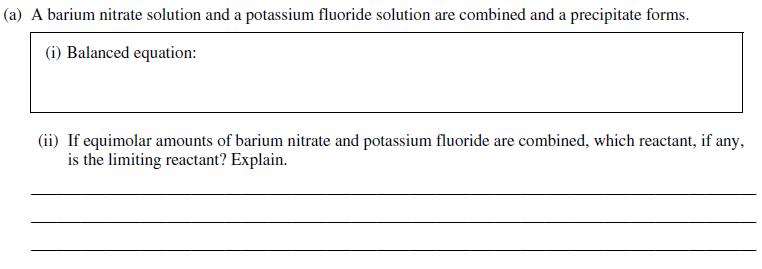
1. What four things must you know to determine if a precipitation reaction will occur? \*summarize!\*
2. Write down the solubility rules. You must remember these as they are not given to you.

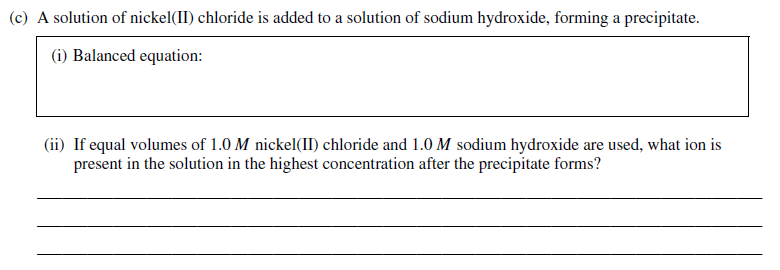
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CHECK FOR UNDERSTANDING



AP PRACTICE QUESTIONS



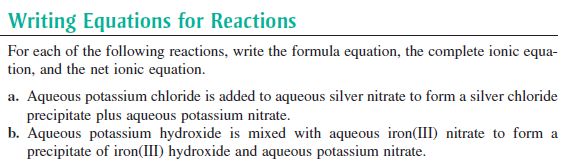


# Describing Reactions in Solutions

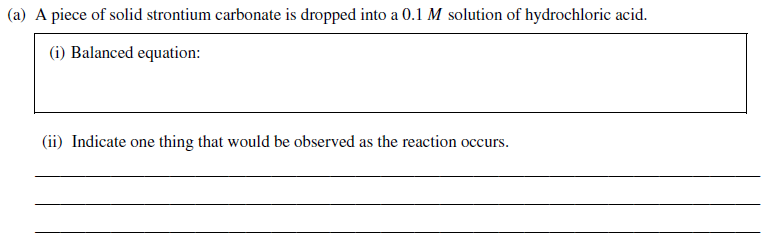
NOTES

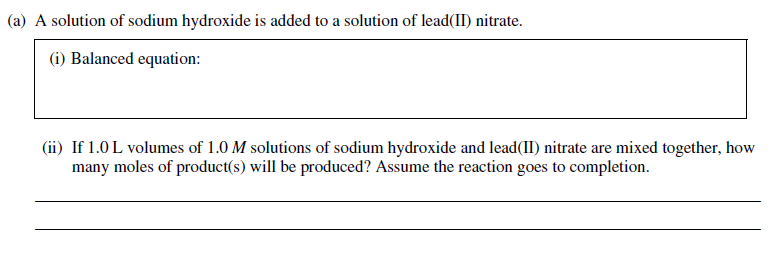
1. Define the following terms:
   1. Strong electrolyte
   2. Formula equation
   3. Complete ionic equation
   4. Net ionic equation
   5. Spectator ions

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AP PRACTICE QUESTIONS





# Acid/Base Reactions

NOTES

1. Compare and contrast Arrehenius’s concept of an acid and a base and Bronsted and Lowry’s definition.
2. Describe how the products in the equation below were formed/determined.

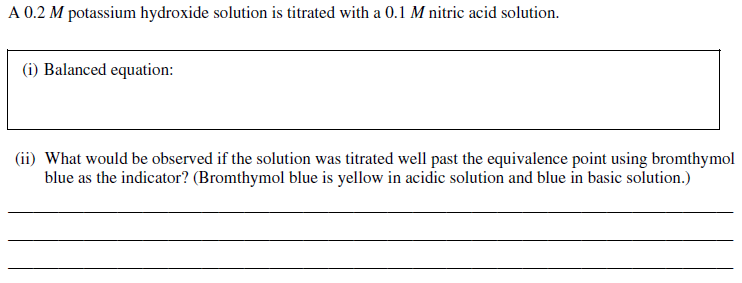


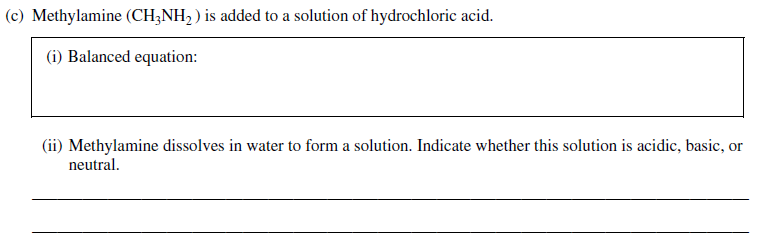
1. What are the steps needed to calculate reactants and products in an acid-base reaction.

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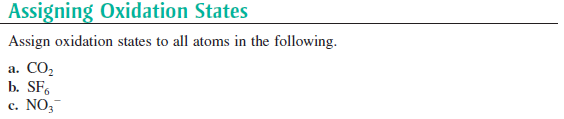
# Redox Reactions

NOTES

1. Define the following terms:
   1. Redox reaction
   2. Reduction
   3. Oxidation
   4. Oxidation number (state)
   5. Oxidizing agent
   6. Reducing agent
2. What are the rules for assigning oxidation states? (Note: there are two slides with the same info on them but in different formats)

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