# **Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ *Regents Chemistry***

## *Unit 11: Nuclear Chemistry*

## *Nuclear Equations*

**Part 1**: Write the complete nuclear equation including mass number and atomic number (charge). Be sure the numbers for the mass and charge add up to the same value on each side.

1. U-238 undergoes alpha decay 2. Pa-234 undergoes beta decay

1. Ra-226 undergoes alpha decay 4. Bi-214 undergoes beta decay

1. Bi-210 undergoes beta decay 6. He-6 undergoes positron decay
2. F-20 undergoes beta decay 8. Po-212 undergoes alpha decay
3. Ne-19 undergoes positron decay 10. U-234 undergoes alpha decay

Look up how each of the following isotopes undergoes decay on Reference Table N and then write the complete decay equation for each isotope.

 

